9. (14 points) Consider the production of lithium nitride from its elements:

$$6 \text{ Li(s)} + \text{N}_2(g)$$
 2 Li₃N(s)

Calculate the percent yield if 3.456 g of lithium produces 4.980 g of product in excess N₂.

Extra Credit (no more than 10 points) Consider the decomposition of sodium azide (NaN₃):

$$2 \text{ NaN}_3(s)$$
 $2 \text{ Na(s)} + 3 \text{ N}_2(g)$

How many grams of sodium azide are required to produce 8.36 L of N_2 gas at 22.5 $^{\circ}\text{C}$ and 0.976 atm?